

CONFECTIONERY PRODUCT

Field of the invention

The present invention relates to premixes for use in the manufacture of aerated
5 confectionery products, especially chilled or frozen confectionery products such
as ice cream.

Background to the invention

A significant market exists for freshly made ice cream, consumed either
10 immediately or soon after preparation, e.g. in a retail outlet/kiosk environment
from a vending machine, or at home. An important element in ice cream
production is the incorporation of air which gives the product proper body and
texture. This incorporated air is termed 'overrun' and is defined as the volume of
ice cream obtained in excess of the volume of the mix, usually expressed as a
15 percentage.

This air is usually incorporated by injection during the freezing and whipping
process. However, the equipment required is costly and therefore only really
commercially viable for large-scale manufacture, or in locations that have a high
20 throughput of customers.

On the other hand, existing means for making ice cream or other aerated
confectionery products on a small scale, such as ice cream makers, do not
provide a satisfactory overrun and/or require time-consuming mechanical
25 techniques to introduce air into the confectionery product.

Accordingly it is an object of the present invention to provide a method of making
aerated confectionery products, such as ice cream, which have the desired level
of overrun, and which can be prepared instantly and hygienically without the use
30 of complex mechanical methods.

Summary of the invention

The present invention is based on the use of dry premixes which comprise a carbon dioxide generating system including an acid and a carbonate. When water is added to the premix, carbon dioxide is produced which introduces the
5 overrun into the product. Other ingredients are included to provide stability to the aerated product over a reasonable period of time, e.g. the time needed to freeze the aerated product in the case of ice cream and other frozen confections.

Carbon dioxide generating systems have been used previously in a number of
10 medicinal, food and beverage products including soluble tablets and ambient confectionery products where they are used to dissolve/disperse ingredients and/or to generate a fizzing sensation. The resulting gas is simply liberated and not retained within the product.

Another problem with carbon dioxide generating systems is that they typically
15 use acid-base chemistry to generate the carbon dioxide by reaction of a weak acid with a carbonate or bicarbonate. However, these systems often lead to a product that is acidic and fizzy which if applied to aerated confectionery products such as ice cream would be undesirable. We have found that by careful choice
20 of the type and amount of acid, it is possible to obtain substantial levels of overrun whilst maintaining the pH above an acceptable level.

Accordingly, the present invention provides a dry confectionery premix for preparing an aerated confectionery product which premix comprises:

- 25 (i) a carbon dioxide generating composition comprising an acid and a carbonate; and
(ii) a stabiliser system;

such that when the premix is mixed with water to give a final solids content of at least about 20 wt%, an aerated confectionery product is formed, in the absence
30 of mechanical aeration, having an overrun of at least about 30% and a pH of greater than about 5.4.

Preferably the premix is a powder.

5 In another aspect, the present invention provides the use of a premix of the invention in a method of preparing a confectionery product having a solids content of at least about 20 wt%, an overrun of at least about 30% and a pH of greater than about 5.4.

10 The present invention also provides a method of preparing a confectionery product which method comprises admixing a premix of the invention with an aqueous liquid to give a final solids content of at least about 20 wt% to form an aerated confectionery product which, in the absence of mechanical aeration, has an overrun of at least about 30% and a pH of greater than about 5.4.

15 In one embodiment, the method further comprises chilling the confectionery product to a temperature of below about 6°C.

In another embodiment, the method further comprises freezing the confectionery product to a temperature of below about -6°C.

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The present invention further provides an aerated confectionery product obtained or obtainable by the method of the invention.

25 In a preferred embodiment, the confectionery product is a chilled or frozen confectionery product, such as ice cream.

Detailed description of the invention

Unless defined otherwise, all technical and scientific terms used herein have the same meaning as commonly understood by one of ordinary skill in the art (e.g. in chilled confectionery/frozen confectionery manufacture, and chemistry).
30 Definitions and descriptions of various terms and techniques used in

chilled/frozen confectionery manufacture are found in Ice Cream, 4th Edition, Arbuckle (1986), Van Nostrand Reinhold Company, New York, NY.

The premixes of the invention are used to make confectionery products, especially chilled or frozen confectionery products such as ice cream, toppings and mousses. Consequently, the bulk of the components of a dry premix of the invention are ingredients typically used in the manufacture of those confectionery products such as a fat source, a protein source, a carbohydrate source, stabilisers, emulsifiers and flavourings. Fat sources include vegetable oil and milk fat. Carbohydrate sources include complex carbohydrate such as starch, and sugars such as sucrose, glucose and lactose. Protein sources include whey and milk solids non-fat. Stabilisers are typically selected from gums, agar, alginates and derivatives thereof, gelatin, pectin, lecithin, sodium carboxymethylcellulose, carrageenan, furcelleran and mixtures thereof.

The term "dry" in the context of the present invention has its usual meaning of 'free from water'. However, it will be appreciated that the premix of the invention may contain a small amount of water, such as less than 5 wt% water, whilst still being essentially 'dry'. The premix is preferably in particulate form, for example in the form of a powder or granules.

The premixes of the invention include a carbon dioxide generating composition comprising an acid and a carbonate. The acid is preferably an acid which is a solid at standard temperature and pressure. The acid is typically a weak acid, preferably an acid with a pKa of greater than about 3. In a preferred embodiment, the acid is selected from a monoprotic acid and a diprotic acid. Preferably the diprotic acid has a first pKa of greater than about 3 and a second pKa of greater than about 4.5, more preferably greater than about 5, 6 or 7. Particularly preferred acids are food grade organic acids such as ascorbic acid, lactic acid, succinic acid or tartaric acid. Mixtures of two or more different acids may be used.

The carbonate can be any suitable carbonate or bicarbonate with any suitable cation, provided that the carbonate is soluble in an aqueous liquid. Examples of suitable cations include metal ions such as potassium and sodium. Mixtures of two or more carbonates may be used.

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The carbonate is present in the premix such that the amount in the final product when made up with an aqueous liquid is from about 0.5 wt% to about 3 wt% of the final product, preferably from about 1 wt% to 2 wt%. Consequently, based on a final solids content of from 20 wt% to 60 wt%, the amount of carbonate present in the premix is from about 0.833 wt% to about 15 wt%. Sufficient carbonate is required to generate the desired overrun. However, the presence of excessive amounts of carbonates can impart an unpleasant taste to the product. The molar ratio of acid to carbonate is selected so as have sufficient acid to react with the carbonate present in the premix and generate carbon dioxide to provide overrun. However, an excess of acid will result in too low a pH. The molar ratio of the amount of acid to base present in the carbon dioxide generating composition is therefore typically from about 1:2 to about 2:1.

The type and amount of acid and carbonate are selected such that when the premix is added to water, the resulting composition has a pH of greater than about 5.4 and the amount of overrun is at least 30%, without using mechanical methods to introduce overrun, i.e. the overrun attributable to the carbon dioxide generating system is at least 30%.

In a preferred embodiment, the pH when measured 30 mins after addition to the dry premix of an aqueous liquid, for example water, is greater than about 5.6, more preferably greater than about 5.8 or 6.0. Preferably, the pH is less than about 8.5, more preferably less than about 8.0 or 7.5. The pH is measured at 10°C.

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The overrun is preferably at least 50%, 70% or 80%, most preferably at least 90% or 100%. The overrun is typically less than 150%. The desired overrun is

preferably maintained for at least 30 mins after addition of the dry premix to an aqueous liquid. In the case of frozen confectionery products which are to be frozen quiescently at -18°C , it is preferred that the desired overrun is maintained for at least 2 hours, more preferably at least 3 or 4 hours to allow sufficient time to freeze the product whilst maintaining the desired overrun.

The overrun is maintained and stabilised by controlling the rheology of the product. If the initial viscosity is too high, this would impede aeration, leading to lower overrun. If the initial viscosity is too low, the gas generated can escape, leading to low overrun.

The rheology of the aerated product can be controlled by choice of the stabilising system. Typical food stabilisers act to viscosify and/or gel the aqueous phase, thus reducing the rate of bubble coalescence, disproportionation and creaming. Food stabilisers are typically selected from gums, agar, alginates and derivatives thereof, gelatin, pectin, lecithin, sodium carboxymethylcellulose, carrageenan, furcelleran. Mixtures of stabilisers can be used. Preferably, the stabiliser is used in combination with a fat based whipping agent, or topping base. These act as a source of fat in the aerated product, which is good for texture and taste. They also act to stabilise the air phase. Examples of commercial fat based whipping agents or topping bases are: Myvatex Mighty Cream and Admul Whippable Fat Powder 2413 (Quest International Ltd), and Instant Topping Base Series DP, e.g. DP73, DP76, (DMV International).

The rheology of the continuous phase produced by pre-mixes can be measured in the following way:

A T.R. Instruments AR 1000-N rheometer is used to measure the sample rheology. A 40 mm acrylic parallel plate geometry (with attached abrasive paper) is used and the gap is set to $1000\text{ }\mu\text{m}$. The correct proportions of powder and water to make 10 g of product are weighed out separately. The acid and base are omitted from the sample, since only the rheology of the mix is required, not

that of the entire foaming product. The powder and water are mixed in a 25 cm³ glass beaker by adding the water to the powder and stirred for one minute by hand with a metal spatula. A small amount of the sample is loaded onto a rheometer. The measurement started two minutes after the powder and water are mixed. A solvent trap is used to prevent the sample drying out during the experiment.

Each formulation is tested with a flow procedure and an oscillation procedure, detailed below. These measurement procedures determine the viscosity of the mix, and the viscous and elastic response of the mix.

Flow Procedure:

Each sample is subjected to a controlled shear rate of 0.5 s⁻¹ for 30 minutes. The temperature is kept constant at 20 °C throughout the experiment and a solvent trap was used. This procedure determines the viscosity of the mix as a function of time.

Oscillation Procedure:

Each sample is subjected to a controlled stress of 0.5 Pa and an oscillation frequency of 1 Hz for 30 minutes. The temperature is kept constant at 20°C throughout the experiment and a solvent trap is used. Data are plotted in terms of the storage (G') and loss (G'') moduli as a function of time, where:

$$G' = G^* \cos \theta$$

$$G'' = G^* \sin \theta$$

θ is the measure of phase angle between applied stress and strain response, and G^* is the instantaneous ratio of stress amplitude to strain amplitude.

From these experiments, we determined that the initial viscosity is preferably from 1 to 100 Pa s (Pascal seconds). After mixing (e.g. 3 minutes), the viscosity has preferably increased and is from 10 to 10000 Pa s. Initially, G' and G'' are

both preferably from 1 to 100 Pa, and G' is less than G'' . After mixing (e.g. 3 minutes), G' and G'' are both preferably from 10 to 1000 Pa, and G' is greater than G'' .

- 5 All viscosity and modulus measurements are taken at standard temperature and pressure.

The premixes of the invention can be used to make confectionery products, especially chilled or frozen confectionery products, by combining the premix with an aqueous liquid such as water or milk. The final wt% solids is typically at least 20 wt%, such as from 20 wt% to 60 wt%, more preferably from 30 wt% to 50 wt%.

After adding the premix to the aqueous liquid, or vice versa, the mixture is typically stirred to ensure that the premix is thoroughly dispersed. Substantial mechanical action is not required: it is generally sufficient to stir the mixture for less than a minute by hand or using an electrically powered stirrer. Such stirring may introduce small amounts of overrun. However, it is not necessary to use mechanical aeration, i.e. mechanical methods that are intended to introduce substantial amounts of gas into the product, such as whipping.

In a preferred embodiment, the premix is not subjected to mechanical aeration.

Nonetheless, the premixes of the invention may be used in conjunction with methods of making confectionery products where the conventional mechanical aeration techniques used do not introduce sufficient overrun. For example, the premixes of the invention may be used in domestic ice cream makers, to enhance the level of overrun to a more acceptable level. In these cases, some of the overrun is introduced by the gas generated composition, and some by mechanical aeration, such as the action of a domestic ice cream making machine. However, the premix must still be capable, in the absence of, mechanical aeration, of generating at least 30% overrun.

The mixture is stirred, or stirred and then allowed to stand, for a sufficient time to allow the reaction between the acid and the carbonate to generate the desired overrun.

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During or after generation of the overrun, the confectionery product may be chilled to a temperature of below about 6°C. If the confectionery product is a frozen confectionery product, then it will be frozen to a temperature of below about -6°C, preferably below -10°C or -15°C. Freezing can be achieved by any
10 suitable means such as in a -18°C freezer. In some embodiments, it may be desirable to use a rapid freezing method such as a brine bath set at a temperature of below -30°C. If used with a domestic ice cream maker, the product will typically freeze during the continuous stirring of the aerated mix.

15 The premixes of the invention may be packaged for retail or home use.

The present invention will now be further described with reference to the following examples, which are illustrative only and non-limiting.

20 **Examples 1 to 6 – Aerated ice cream**

Materials and methods (Examples 1 to 6)

100g of powder was made up by blending the individual dry powder ingredients together (see table 1 for list of ingredients). 175 g of cold water (10°C or less)
25 was prepared. Half of this was added to the dry ingredients and mixed with a spoon. Once the ingredients were dispersed, the remainder of the cold water was added with further stirring until a homogeneous foamed mix was produced. The total time from initial water addition to completion is about 1 minute.

30 The pH and overrun of the aerated mix was measured immediately after formation using standard methods.

The aerated mix was then portioned into approximately 5 small plastic containers (roughly 30 to 50 g in each), and then quiescently frozen in a -25°C room.

5 After 48 hours, the overrun of the ice cream was measured using the Archimedes' principle as described below, and the ice cream was melted and the pH measured at 10°C.

For each example, the process was repeated at least 3 times.

10

Determining overrun of a finished product

The density of a finished ice cream (or other aerated ice confection) product can also be estimated by making use of the Archimedes' principle as described in
15 "A-level Physics", Third Edition, by R. Muncaster, Pub. Stanley Thornes Ltd., Cheltenham, 1989.

First a sample of ice cream is weighed in air to determine its mass. Then the volume of the same sample is determined using the Archimedes' principle as
20 described below. The sample of ice cream is held carefully in a beaker of chilled water just below the surface of the water by a fork (or a knife) inserted into the end of the product. The beaker is placed on a balance throughout the experiment and the increase in weight on immersing the product is recorded. By Archimedes' principle, the increase in weight is equal to the upthrust and hence
25 weight of water displaced. Taking the density of water as 1 gcm^{-3} , the weight of water displaced is used to determine the volume of water displaced and thus the volume of ice cream immersed in the beaker. From the mass and volume of the product, the density of the ice cream can be calculated. A minimum of three repeat measurements is taken.

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The density of the unaerated mix can either be assumed to be 1.12 g/cm^3 or can be estimated by melting the ice cream until the air-phase is lost and then

determining the density in an overrun cup at 4°C as described above. With a knowledge of the density of both unaerated mix and aerated ice cream, the overrun can be calculated using the following equation:

$$5 \quad \text{overrun \%} = \frac{\text{density of mix} - \text{density of ice cream}}{\text{density of ice cream}} \times 100$$

Table 1 – formulation ingredients

	Example 1		Example 2		Example 3	
Ingredient	wt.% (1)	wt.% (2)	wt.% (1)	wt.% (2)	wt.% (1)	wt.% (2)
Skim Milk Powder	6.6	2.400	7.6	2.764	6.6	2.400
Sucrose	46	16.727	47.86	17.404	46	16.727
Myvatex Mighty Cream	38	13.818	38	13.818	38	13.818
Xanthan Gum	1	0.364	1	0.364	1	0.364
k-Carrageenan	0.11	0.040	0	0.000	0.11	0.040
Vanillin	0.04	0.015	0.04	0.015	0.04	0.015
Acid	5.5	2.000	2.75	1.000	5.5	2.000
Base	2.75	1.000	2.75	1.000	2.75	1.000

	Example 4		Example 5		Example 6	
Ingredient	wt.% (1)	wt.% (2)	wt.% (1)	wt.% (2)	wt.% (1)	wt.% (2)
Skim Milk Powder	6.6	2.400	7.725	2.809	6.6	2.400
Sucrose	46	16.727	49	17.818	46	16.727
Myvatex Mighty Cream*	38	13.818	38	13.818	38	13.818
Xanthan Gum	1	0.364	1	0.364	1	0.364
k-Carrageenan	0.11	0.040	0.11	0.040	0.11	0.040
Vanillin	0.04	0.015	0.04	0.015	0.04	0.015
Acid	5.5	2.000	1.375	0.500	5.5	2.000
Base	2.75	1.000	2.75	1.000	2.75	1.000

Key:

5 wt%(1) = weight percentage of ingredient in the dry powder.

wt%(2) = weight percentage of ingredient in the hydrated product.

* obtained from Quest International, the 38 wt%(1) consisting of, according to the manufacturer, 24.7 wt% fat minimum, 3.04 wt% protein (whey) minimum, the remainder being made up from mono-/di-glycerides of fatty acids, guar gum, agar and locust bean gum.

Table 2 – amount of acid and base

	Ascorbic Acid		Citric Acid		Sodium Bicarbonate		Calcium Carbonate	
	A	B	A	B	A	B	A	B
Example 1: Conc. / wt %	5.5	2	-	-	2.75	1	-	-
Example 2: Conc. / wt %	2.75	1	-	-	2.75	1	-	-
Example 3: Conc. / wt %	-	-	5.5	2	2.75	1	-	-
Example 4: Conc. / wt %	-	-	2.75	1	2.75	1	-	-
Example 5: Conc. / wt %	-	-	1.375	0.5	2.75	1	-	-
Example 6: Conc. / wt %	5.5	2	-	-	-	-	2.75	1

A = wt.% in base powder

B = wt.% when aqueous solution added

Results

5

Table 3 – summary of results

	Overrun / %		pH	
	X	Y	X	Y
Example 1:	101	101	5.9	6.6
Example 2:	61	70	6.4	7.6
Example 3:	82	114	3.9	4.3
Example 4:	111	79	5.5	5.5
Example 5:	76	74	6.6	8.3
Example 6:	41	66	4.1	4.5

X = After aeration

Y = After freezing

Discussion of examples 1 to 6

Example 1: 2% Ascorbic acid and 1% sodium bicarbonate.

5 A good level of overrun is obtained after aeration (100%) that is maintained through freezing. The pH of the final product is also well above 5.4, ensuring that the ice cream does not taste acidic. The pH measured after melting is greater than that measured immediately after mixing since over time the carbon dioxide will evaporate from solution, reducing the amount of hydrogen ions in
10 solution.

Example 2: 1% Ascorbic acid and 1% sodium bicarbonate.

As the amount of acid is reduced, the level of overrun also decreases and the pH remains high. The overrun in this example is lower than would be preferred,
15 indicating that when using 1% of sodium bicarbonate in the final product, at least 1.5 to 2% ascorbic acid is required to obtain the preferred overrun.

Example 3: 2% Citric acid and 1% sodium bicarbonate.

A good amount of overrun is obtained, measured as 114% after freezing.
20 However, the pH is well below the required threshold of 5.4, leading to a very acidic tasting and fizzy ice cream product which is undesirable. The pH measured after melting is similar to that measured immediately after mixing since although carbon dioxide will evaporate from solution, the citric acid has excess protons, keeping the pH low.

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Example 4: 1% Citric acid and 1% sodium bicarbonate.

In this case, after freezing and overrun of 79% is achieved. This is an acceptable level of gaseous phase. However, although the pH is also above the required threshold (measured as 5.5 after freezing), this choice of acid and base

is not ideal since if a greater degree of overrun was desired, then further addition of citric acid would lead to a pH less than 5.4. Also, small discrepancies in weighing ingredients could lead to a similar result.

5 **Example 5: 0.5% Citric acid and 1% sodium bicarbonate.**

An acceptable pH is obtained when using this mix, although the overrun in the final product is not as high as preferred. Examples 3 to 5 show that small changes in citric acid concentration have a greater effect on both pH and overrun relative to use of ascorbic acid. Use of 1% sodium bicarbonate with citric acid
10 does not achieve the most preferred overrun and pH of the final aerated product.

Example 6: 2% Ascorbic acid and 1% Calcium Carbonate.

In this example, the overrun is much less than that preferred. This is because calcium carbonate has a low solubility in water compared to sodium bicarbonate.
15 Therefore, the reaction is much slower and does not go through to completion before the product is frozen. Further, the pH remains low since the acid is not all used in reaction. It is clear that choice of base is important to get the rate of reaction optimised.

20 **Example 7 - Modelled data**

To investigate further the findings shown in examples 1 to 6, a model was developed to enable us to predict the behaviour of different acids in solution with respect to their reaction with a base; the output of the model being the resulting amount of carbon dioxide which is liberated by the reaction and the solution pH.
25 The model assumes the solution is water, which is saturated with carbon dioxide gas. Any carbon dioxide that evolves is therefore liberated immediately as gas, as opposed to dissolving in solution.

pH values and carbon dioxide volumes as a function of acid type and
30 concentration are shown in figure 1. It is clear that for ascorbic acid, as a

function of concentration, the curves for volume of carbon dioxide produced and the pH are shallow compared to the other di-protic and tri-protic acids plotted.

Citric acid has the steepest curve, which means that small changes in acid concentration have a greater effect on pH and volume of carbon dioxide. The curve is particularly steep around the final pH values of 5 and 6, i.e. very small concentration changes will result in a significant pH change. Such dramatic changes are less preferred since control of these parameters is more difficult, and it is therefore more difficult to generate a product that falls inside the desired pH and overrun window, especially when the product is for home use or use by unskilled retailers.

Overall, the pH and volume of carbon dioxide curves for ascorbic acid have a smaller gradient than that of citric acid. Ascorbic acid is a preferred acid since it exhibits both a low pKa and a high pKa (4.17 and 11.57). Di- and tri-protic acids that have pKa values that are all low (< 5) are less preferred since they will exhibit steeper pH and volume carbon dioxide curves.

Table 4: pKa values

20

	Acid	MW/g mol ⁻¹	pKa
Monoprotic	Acetic	60.05	4.756
	Lactic	90.08	3.863
	Hydrochloric	36	-7
	Gluconic	196	3.86
	Nitric	63	-1.444
	Mandelic	152.15	3.42
	Crotonic	86.09	4.69
	Pyruvic	88.06	2.5

Diprotic	Glycine	75.07	2.35, 9.77
	Succinic	118.09	4.207, 5.636
	Oxalic	126.07	1.268, 4.282
	Malic	134.09	3.4, 5.1
	Tartaric	150.09	3.2, 4.8
	Ascorbic	176.13	4.17, 11.57
Triprotic	Citric	192.13	3.13, 4.757, 5.602

The various features and embodiments of the present invention, referred to in individual sections above apply, as appropriate, to other sections, *mutatis mutandis*. Consequently features specified in one section may be combined with features specified in other sections, as appropriate.

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All publications mentioned in the above specification are herein incorporated by reference. Various modifications and variations of the described methods and products of the invention will be apparent to those skilled in the art without departing from the scope of the invention. Although the invention has been described in connection with specific preferred embodiments, it should be understood that the invention as claimed should not be unduly limited to such specific embodiments. Indeed, various modifications of the described modes for carrying out the invention which are apparent to those skilled in the relevant fields are intended to be within the scope of the following claims.

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